

TESTING PRIOR KNOWLEDGE

Title: CORROSION RATES

Measuring the rate of iron corrosion

Choose the correct answer.

1. Dissolve 1.404 g $\text{Fe}(\text{NH}_4)_2(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$ in a 1000 cm^3 volumetric flask. What will be the Fe^{2+} content in 1 cm^3 of solution?

- a. 1.8 mg
- b. 0.2 mg
- c. 0.1 g
- d. 0.5 mg

2. If the transmission of solution is 50%, the absorbance is...

- a. 1.500
- b. 0.500
- c. 5.72
- d. 0.301

3. Aqueous corrosion of iron requires:

- a. H_2O
- b. O_2
- c. both
- d. air

4. What mass of iron is present in 0.5267 g of ammonium-iron(II) sulfate-6-water?

- a. 10 mg
- b. 75 mg
- c. 100 mg
- d. 750 mg

Note: answers are given in red. These need to be removed before the test is used.

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